

EOS Mixing Rule Parameters

GENPAR

Fitting of Equation of State Mixing Rule Parameters
for Flash and VLE Calculation

DDBSP - Dortmund Data Bank Software Package



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Introduction

Generate EOS Mixing Rule Parameters (GenPar) is the fit program for equation of states mixing rule parameters which can be used to describe phase equilibria up to the critical point of a mixture. This is the advantage over g^E -models which can be used only at moderate pressures and temperatures not too far away from normal boiling points.

The *GenPar* program uses binary data sets from the Dortmund Data Bank parts VLE and HPV which both contain vapor-liquid equilibrium data. HPV contains mostly data where the normal boiling points of at least one component is below 273 K whereas the VLE data bank contains all other data sets.

GenPar can fit parameters for the equations of state

- Soave-Redlich-Kwong
- Peng-Robinson
- Redlich-Kwong

The EOS Redlich-Kwong is not available until now but the routines are prepared.

The supported mixing rules are

- quadratic mixing rules
- four g^E mixing rules:
 - Vidal
 - Tochigi
 - Michelsen
 - PSRK
- Panagiotopoulos/Reid

Searching for Components and Experimental Data Sets

Figure 1 shows the start screen of *Generate EOS Mixing Rule Parameters*.

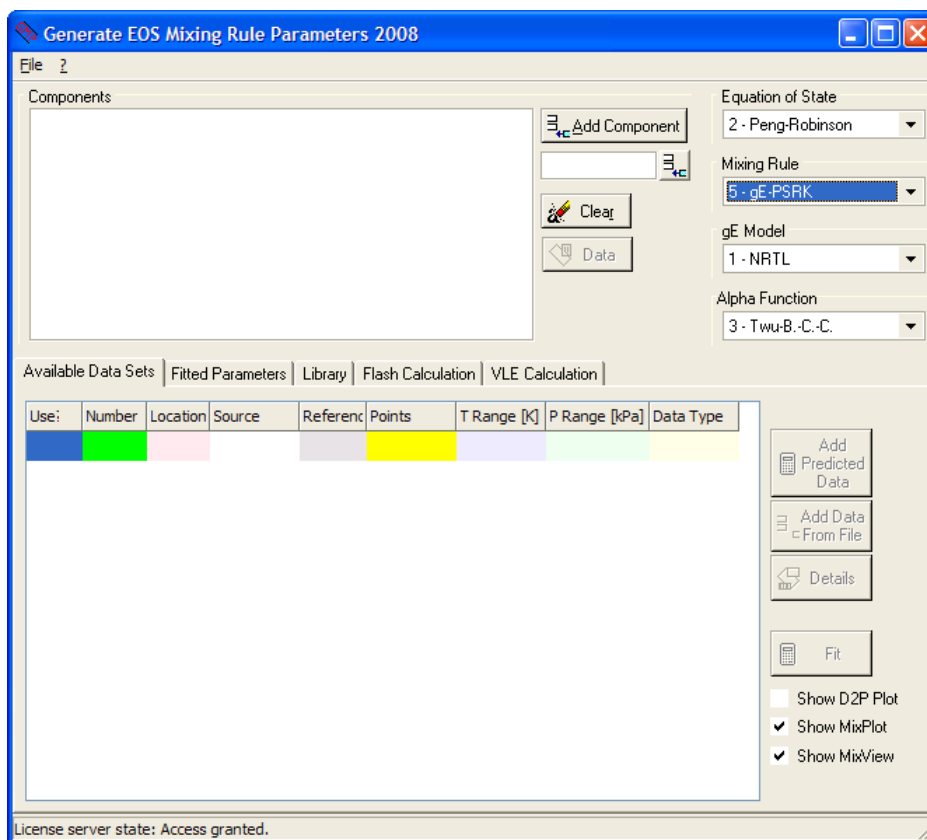
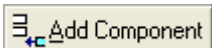



Figure 1: Start Screen

The *Components* group box contains an initially empty list plus these controls:

- 

This button invokes the *Compound Selection* tool (see Figure 2).

- 

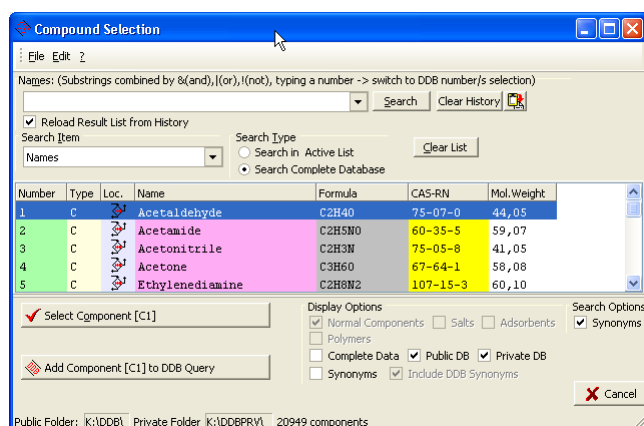
If the DDB number of a component is known it can be typed in directly.

- 

This button clears the current component list

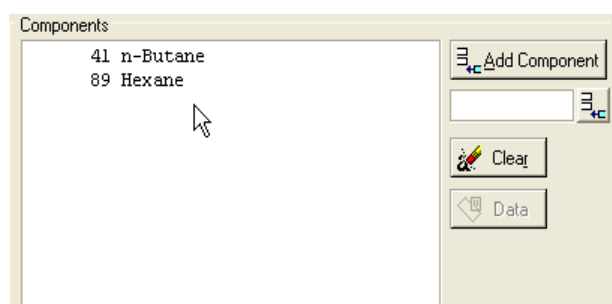
- 

If a single component is selected this button displays the basic data bank entries like name, CAS registry number, formula etc.



The *Compound Selection* dialog is described in detail in the “Component Management” documentation.

Figure 2: Component Selection



After selecting at least two components (see Figure 3) the VLE and HPV data banks are searched for the binary data sets.

Figure 3: Component List

The *Available Data Sets* card shows the different binary systems and the data sets which have been found in both data banks.

Use?	Number	Location	Source	Reference	Points	T Range [K]	P Range [kPa]	Data Type
<input checked="" type="checkbox"/> Yes	582	0 - Public DB	0 - HPV	129	9	358 - 426	1013	Txy(P)
<input checked="" type="checkbox"/> Yes	583	0 - Public DB	0 - HPV	129	9	378 - 450	1519	Txy(P)
<input checked="" type="checkbox"/> Yes	584	0 - Public DB	0 - HPV	129	9	394 - 468	2026	Txy(P)
<input checked="" type="checkbox"/> Yes	585	0 - Public DB	0 - HPV	129	9	408 - 484	2533	Txy(P)
<input checked="" type="checkbox"/> Yes	586	0 - Public DB	0 - HPV	129	8	426 - 493	2938	Txy(P)
<input checked="" type="checkbox"/> Yes	587	0 - Public DB	0 - HPV	129	7	436 - 492	3394	Txy(P)
<input checked="" type="checkbox"/> Yes	588	0 - Public DB	0 - HPV	129	4	442 - 475	3698	Txy(P)
<input type="checkbox"/> No	572	0 - Public DB	0 - HPV	572	2	396 - 416	861	Tx(P)
<input type="checkbox"/> No	493	0 - Public DB	0 - HPV	572	4	375 - 426	1034	Tx(P)
<input type="checkbox"/> No	1025	0 - Public DB	0 - HPV	572	4	378 - 435	1206	Tx(P)

Figure 4: Available Data Sets, Single System

If the component lists gets longer (see Figure 5) the list of binary component pairs gets longer (see Figure 6).

41 n-Butane
 89 Hexane
 161 Toluene
 1056 Nitrogen
 237 Propane

Figure 5: Component List

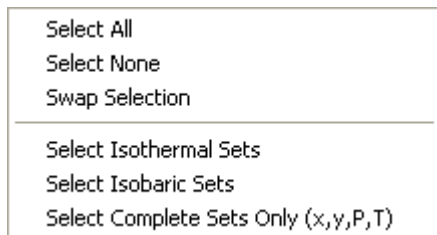
The data set list contains the following columns:

Use?	Number	Location	Source	Reference	Points	T Range [K]	P Range [kPa]	Data Type
<input checked="" type="checkbox"/>	432	0 - Public DB	0 - HPV	90	6	143	1379 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	433	0 - Public DB	0 - HPV	90	7	173	6 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	434	0 - Public DB	0 - HPV	90	7	198	22 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	435	0 - Public DB	0 - HPV	90	7	223	70 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	436	0 - Public DB	0 - HPV	90	7	248	205 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	437	0 - Public DB	0 - HPV	90	7	273	468 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	438	0 - Public DB	0 - HPV	90	7	298	930 - 13790	Pxy(T)
<input checked="" type="checkbox"/>	439	0 - Public DB	0 - HPV	90	5	323	1689 - 12411	Pxy(T)
<input checked="" type="checkbox"/>	440	0 - Public DB	0 - HPV	90	5	333	2137 - 11032	Pxy(T)
<input checked="" type="checkbox"/>	441	0 - Public DB	0 - HPV	90	5	343	2551 - 8273	Pxy(T)

Figure 6: Available Data Sets, Many Systems

1. **Use?**
These check boxes determine if the data set shall be used for fitting
2. **Number**
Data set number
3. **Location**
Location of the data bank from which the data set has been loaded (0 stands for a public DDB from DDBST GmbH, 1 stands for a private data bank from a user)
4. **Source**
0 if the data sets has been read from the HPV data bank and a one for the VLE data bank
5. **Reference**
Article number
6. **Points**
Number of points of the data set
7. **T Range [K]**
Temperature range of the data set
8. **P Range [kPa]**
Pressure range of the data set.
9. **Data Type**
Indicates the data columns of a data set. Column names in parenthesis denotes constant values. So "Pxy(T)" means: pressure, liquid and vapor composition at constant temperature is specified.

The grid has a context menu for modifying selections.



The entry ‘Swap Selection’ selects previously unused sets and deselects previously checked sets.


The other entries allow to create sublists with matching conditions.

Add Predicted Data Sets

If the DDB contains no or insufficient data sets for a binary system it is possible to create data sets by predictive methods.

The dialog allows to select the predictive methods and enter the conditions.

This dialog is also part of the mixture properties data bank management program and is described in that documentation in detail.

The  button produces a predicted VLE data set and returns this data set to *Generate EOS Mixing Rule Parameters*.

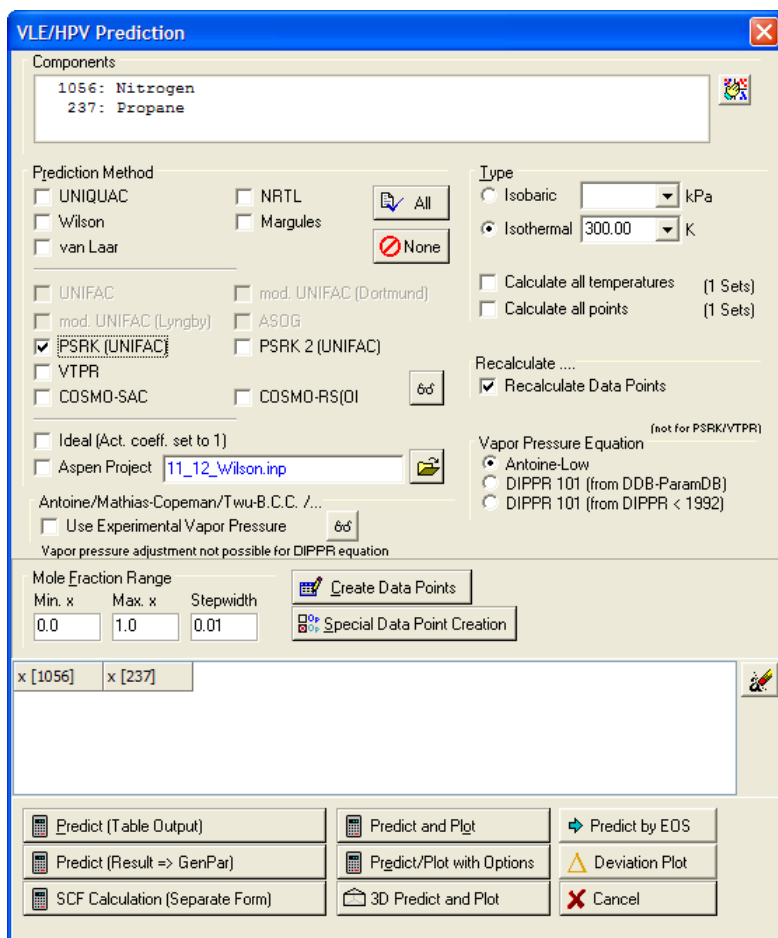


Figure 7: Prediction of Vapor-liquid Equilibria

Use?	Number	Location	Source	Reference	Points	T Range [K]	P Range [kPa]	Data T
<input checked="" type="checkbox"/> Yes	1786	0 - Public DB	1 - VLE	811	6	362 - 422	1013	Txy(P)
<input type="checkbox"/> Yes	4227	0 - Public DB	1 - VLE	1277	12	293	16 - 206	Px(T)
<input type="checkbox"/> Yes	4228	0 - Public DB	1 - VLE	1277	12	283	10 - 147	Px(T)
<input type="checkbox"/> Yes	4229	0 - Public DB	1 - VLE	1277	12	273	6 - 103	Px(T)
<input type="checkbox"/> Yes	4230	0 - Public DB	1 - VLE	1277	12	263	3 - 69	Px(T)
<input type="checkbox"/> Yes	4231	0 - Public DB	1 - VLE	1277	12	253	1 - 45	Px(T)
<input checked="" type="checkbox"/> Yes	0	205 - Pred.	0 - HPV	0	51	300	21 - 257	Pxy(T)

The location number is 205 indicating that the PSRK model created the data set. The “Source” and “Reference” field has to be ignored.

Hint

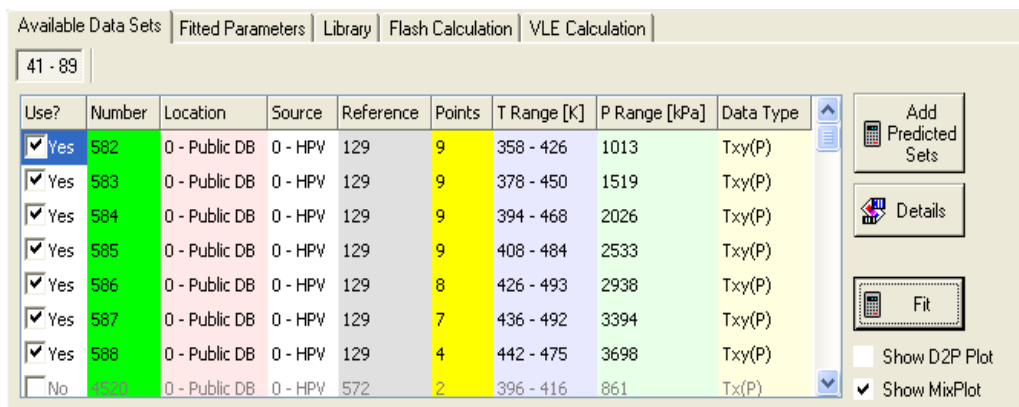
Many models are restricted to moderate pressures and temperatures near the normal boiling point. Group contribution methods need group assignments and group contributions. g^E models need g^E parameters. There are many possibilities that a calculation can fail.

Add Data From File

If data are available from other sources or have altered outside the DDB system or saved for any reason it is possible to add data from files. The data have to be in “.qr” format that the Dortmund Data Bank programs also reads and writes.

Fitting Mixing Rule Parameters

First select the *Available Data Sets* card, the select a binary system (if there are more than two components in the system). When some binary data sets have been selected to be used in the fit process (the “Use?” check boxes are checked) the *Fit* button becomes active:



The step before or after selecting the components and loading the data sets is the selection of the equation of state, the mixing rule, and the α -function.

If a g^E mixing rule has been selected the g^E model has to be specified additionally.

1 - Quadratic			
2 - gE-Vidal		1- NRTL	
3 - gE-Tochigi		2 - Wilson	
4 - gE-Michelsen		3 - UNIQUAC	1 - Mathias-Copeman
5 - gE-PSRK	1 - Soave-Redlich-Kwong	4 - NRTL (Vidal)	2 - Stryjek-Vera
6 - Pan./Reid	2 - Peng-Robinson	5 - UNIQUAC (res)	3 - Twu-B.-C.-C.
	3 - Redlich-Kwong		4 - Acentric Factor

In case of the mixing rule by *Tochigi* the g^E model is fixed to *UNIQUAC (res)*.

Now the *Fit* button can be pressed. An option dialog will be displayed (see Figure 8). It is possible to use temperature independent parameters (*none*), *linear* or *quadratic* temperature dependent parameters.

The default starting values (“0”s) for the parameters can be modified. The alpha value (NRTL) will stay constant during the fit process.

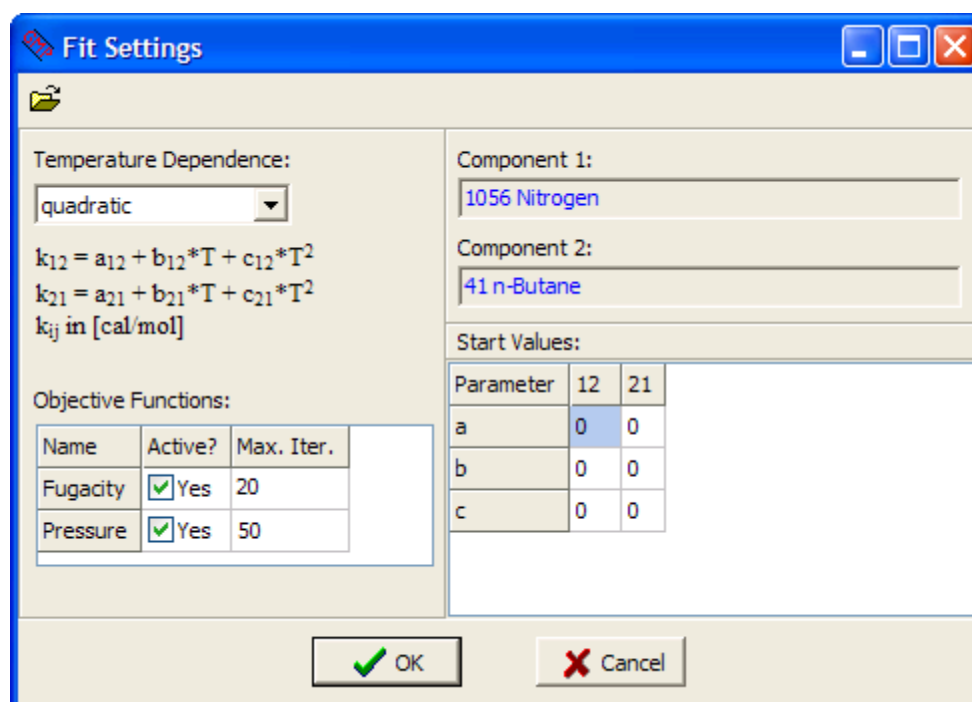


Figure 8: Options for parameter fitting



It is also possible to load starting values from EOS project files (*.eosproj), fit input files (*.eoi) or fit output files (*.out). All these files are text files with parameter lines beginning with “KIJ”.

The optimization process can use up to two objective functions:

- minimize *fugacity* differences
- minimize *pressure* differences.

At least one objective function has to be selected. If both are selected the sum of both objective functions will be minimized.

The optimization process takes a short while. A window with the optimization results (text) is displayed. If the calculation process detects a miscibility gap it is possible to update the VLE calculation results (see Figure 9).

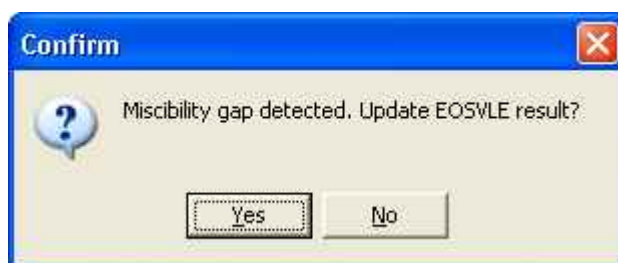


Figure 9: Miscibility gap detection dialog

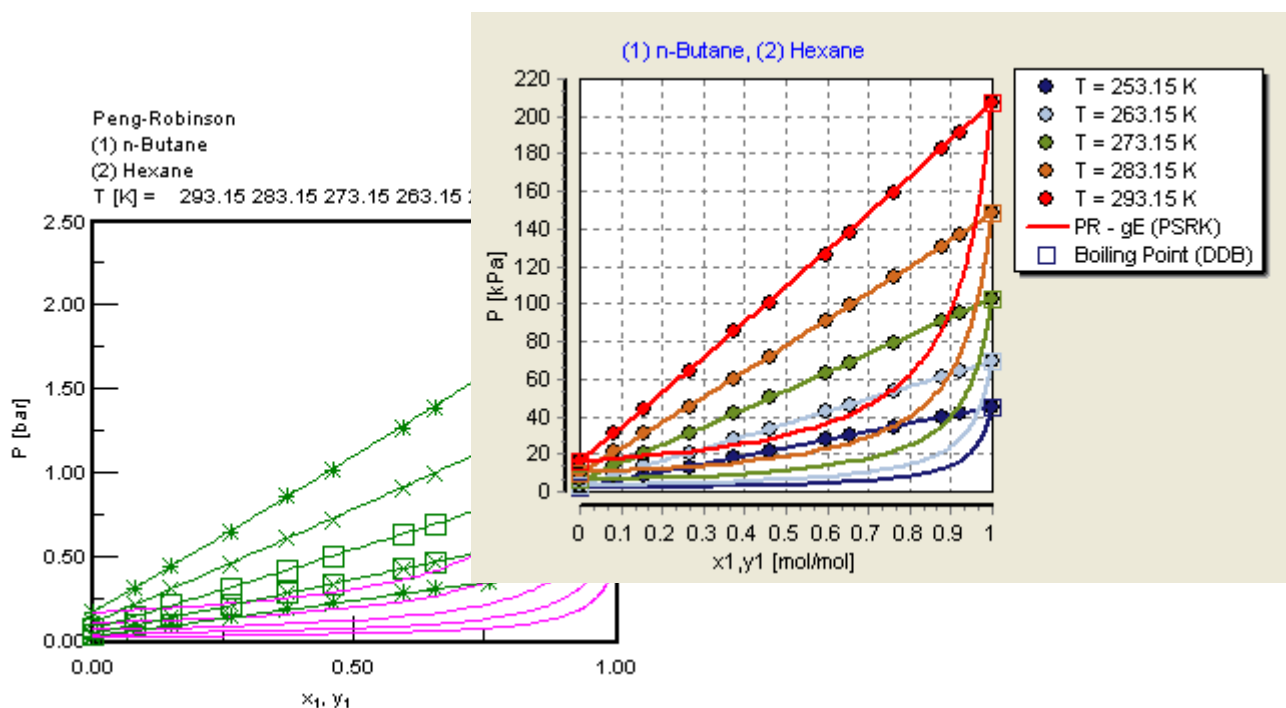
There are up to three result cards:

EOSOPTLP - Fit Log | EOSVLE - Fit Result Comparison | EOSVLE - Miscibility Gap Information

1. The fit result created by EOSOPTLP. This text has the following sections:
 - a. Duplicate of the EOSOPTLP command file (“INPUT DATA”)
 - b. The list of data sets used for the fit (“DATA SETS IN FILE”)
 - c. The specified EOS and mixing rule (“EQUATION OF STATE”)
 - d. Existing mixing rule parameters (“INTERACTION PARAMETERS...”, always 0.)
 - e. A list of modified acentric factors or the first Mathias-Copeman parameter for each data set used in the fit. This value is adjusted to pure component vapor pressures of the single sets (“SYSTEM: ??? C1(2)=...”).
 - f. The minimizing fugacity differences progress report (“MINIMIZING FUGACITY DIFFERENCE”)
 - g. The minimizing relative pressure deviation (“MINIMIZING RELATIVE PRESSURE DEVIATIONS”)
 - h. The resulting interaction parameters (= mixing rule parameters, “INTERACTION PARAMETERS FOR ...”)
 - i. Deviation tables for all used data sets.
2. The second card contains tables of VLE data for all temperatures or pressures given in the used data sets. These VLE are calculated with the fitted parameters
3. In the cases where a data set contains a miscibility gap a third card gives information on the found LLE.

If *Show MixPlot* is selected the calculated VLE data together with the experimental data sets are displayed using the standard DDBSP plot window.





If *Show D2P Plot* is selected the D2P plot output is used to display the experimental VLE together with the calculated sets.

The fitted parameter sets are stored in *GenPar*.

Available Data Sets		Fitted Parameters	Library	Flash Calculation	VLE Calculation			
Comp't 1	Comp't 2	EOS	Mixing Rule	gE Model	a12	b12	c12	
41	89	2 - Peng-Robinson	5 - gE (PSRK)	1 - NRTL	145.6563753	0.08087628127	0	Save as Project
1050	41	2 - Peng-Robinson	5 - gE (PSRK)	1 - NRTL	-34.3090756	3.117049372	0.00166868	Save Set in Database
1050	89	2 - Peng-Robinson	5 - gE (PSRK)	1 - NRTL	2215.617637	-11.15781455	0.00226942	Clear

All parameters sets which are created in one program run (a *session*) are collected in that grid. Only one parameter set per component pair is stored here. A new fit overwrites any previously fitted set.

The sets can be stored either in an EOS project file or in a so called library (DDB parameter database). The project files are used to stored a list of components together with their pure component properties and the mixing rule parameters for a specific EOS and mixing rule and g^E model. To be backwards compatible it is possible to load the EPF files which were used in previous versions of the *Generate EOS Mixing Rule Parameters* program. The format of EPF files are described in the appendix.

The fitted parameters can be deleted completely by clicking the *Clear* button or removed line by line by clicking the *Remove Line* button.

Working with the Library

The library is essentially a list of parameter sets. It's divided in two parts, a public list which is distributed by DDBST GmbH and a second, private list which can be used by our customers.

Parameter sets cannot be edited directly. Adding data sets is possible only for the private library from the *Fitted Parameters* list (see section Fitting Mixing Rule Parameters) and complete lines (which are parameter sets) can be deleted by the *Remove* button.

		No.	Comp't 1	Comp't 2	EOS	Mixing Rule	PE Model	g12	g21	alpha	Tmin	
Public Lib		4435	41	174	1	1	0	-0.2163	0	0	277	Remove
		4436	174	237	1	1	0	-0.2768	0	0	277	
Private Lib		4437	41	1050	1	1	0	0.1525	0	0	227	
		4438	60	1050	1	1	0	0.1441	0	0	277	
		4439	89	1050	1	1	0	0.1326	0	0	293	
		4440	91	1050	1	1	0	0.1361	0	0	283	
		4441	94	1050	1	1	0	0.1348	0	0	277	
		4442	128	1050	1	1	0	0.154	0	0	283	

Flash Calculation

Generate EOS Mixing Rule Parameters allows to invoke a two-phase and a three-phase flash calculation.

The two-phase flash calculates composition of the vapor and the liquid phase without considering miscibility gaps whereas the three phase flash tries to find the liquid-liquid equilibrium if it exists.

The flash is calculated either isothermal within a pressure range or isobaric within a temperature range.

Temperature: 323.15 K
 Pressure Range [bar]: 1.0, 2.0, 0.1
 Constant Temperature
 Constant Pressure

EOS, MR, and gE-model settings, see here

Component	Mole Fractions
41	0.23
89	0.64
1050	0.13

Type: Mole Fractions

2 Phase Flash
3 Phase Flash
Create Grid
Clear Grid

Current Sum: 1

Comp't 1	Comp't 2	a12	b12	c12	a21	b21	c21
41	89	145.6563753	0.08087628127	0	-151.8775621	-0.163167687	0
1050	41	-34.3090756	3.117049372	0.001668682746	1416.63341	-5.838123026	-0.0
1050	89	2215.617637	-11.15781455	0.002269420481	-4527.749766	22.45724966	-0.0

Kij

The global composition (“Feed”) can be specified as mole fraction, mole percent, moles or masses (kilograms), weight fractions or weight percent.

- Fill by Newly Fitted Parameters (1.)
- Fill by Private Library Parameters (2.)
- Fill by Public Library Parameters (3.)
- Fill by 1., 2., 3.

Clear All

The parameters used for the flash calculation can be obtained from new fitted parameter sets, from the private or public parameter library or by all three sources. The import can be invoked from a context menu of the parameter grid. The context menu is also shown by the button **Kij**.

Another source is a project file which can be loaded from the *File* menu. And additionally it is possible to type in the parameters by hand.

After completing the input of the global composition and the parameters the flash calculation can be started.

Two-Phase Flash Results

A two-phase flash gives a text result and a plot result.

The text output contains the following information:

```
EQUATION OF STATE: PENG-ROBINSON
MIXING RULE: PSRK
ALPHA FUNCTION: TWU-BLUCK-CUNNINGHAM-COON
GE-MODEL: NRTL
```

```
I TC[K] PC[BAR] OMEGA C1 ...
1 304.20 73.76 0.22520 0.82105 ...
2 425.20 38.00 0.19300 0.40119 ...
3 507.40 30.14 0.29750 0.38727 ...
```

INTERACTION PARAMETERS FOR GE(NRTL) MIXING RULE:

```
I J AIJ AJI BIJ ...
1 2 -34.309 1416.6 3.1170 ...
1 3 2215.6 -4527.7 -11.158 ...
2 3 145.66 -151.88 0.80876E-01 ...
```

ISOBARIC IDEAL GAS HEAT CAPACITY COEFFICIENTS:

```
I Eq A1 A2 A3 ...
1 9 29370. 34540. 1428.0 ...
2 9 71340. 0.24300E+06 1630.0 ...
3 9 0.10440E+06 0.35230E+06 1694.6 ...
```

DEFINITIONS USED IN THIS PROGRAM:

```
IDEAL GAS ENTHALPY = 0 AT T0=298.15 K
IDEAL GAS ENTROPY = 0 AT T0=298.15 K AND
P0=1 ATM
THE FOLLOWING PENELOUX CONSTANTS [L/MOL] ARE
USED FOR VOLUME CORRECTIONS:
-0.00182 -0.00492 -0.00212
```

The first lines specify the equation of state settings together with the used vapor pressure equation for pure components (alpha function).

The used pure component properties follow.

Then the mixing rule parameters are shown.

Pure component heat capacity parameters.

These definitions describe some basic definitions for the used models.

Peneloux constants.

The following line are repeated for every new pressure (isothermal calculations) or new temperatures (isobaric calculations).

TEMPERATURE:	323.15 K						
PRESSURE :	1.0000 BAR						
I	Z (I)	F*Z	X (I)	L*X	Y (I)	V*Y	K (I)
1	0.13000	0.13000	0.00159	0.00048	0.18629	0.12952	117.518
2	0.23000	0.23000	0.06697	0.02041	0.30146	0.20959	4.501
3	0.64000	0.64000	0.93144	0.28387	0.51225	0.35613	0.550
N [MOL]	1.00000		0.30476		0.69524		
M [G]	74.24		25.67		48.57		
V [L/MOL]			0.1327		26.11		
Z=PV/RT			0.0049		0.9716		
V [L/MOL]	(Corrected)		0.1351		26.11		
Rho [g/ccm]	(Corrected)		0.624		0.003		
H [J/mol]			-26030.0		2619.9		TOTAL: -6111.4
S [J/mol/K]			-75.90		8.80		TOTAL: -17.01

Abbreviations - component specific values

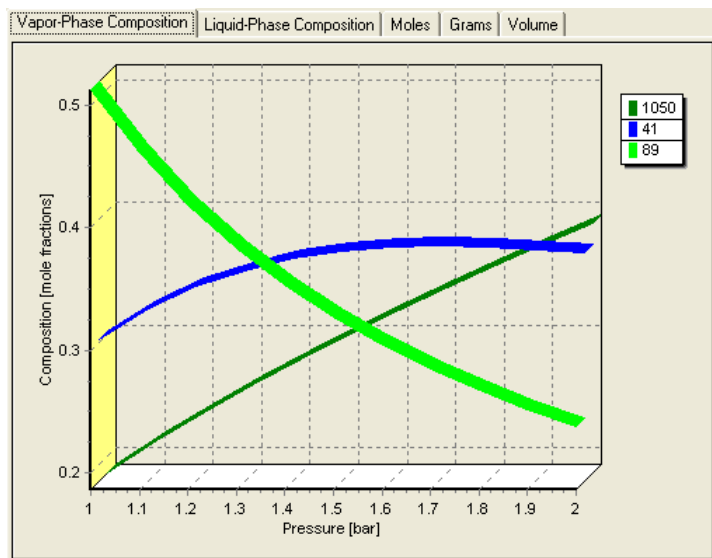
- I Component (numbering order from top to bottom as in the Component List)
- Z(I) Total number of moles of component I (normalized to 1 mole of total feed)
- X(I) Mole fraction of component I in the liquid phase
- Y(I) Mole fraction of component I in the vapor phase
- K(I) $Y(I)/X(I)$
- F Total feed (1 mole)
- L Total number of moles in the liquid phase
- V Total number of moles in the vapor phase

Abbreviation – system specific values

- N [MOL] Moles (complete, in liquid phase, in vapor phase)
- M [G] Masses (complete, in liquid phase, in vapor phase)
- V [L/MOL] Volume (of liquid phase, of vapor phase)
- Z=Pv/RT Compressibility (of liquid phase, of vapor phase)
- Rho [g/cm] Density (of liquid phase, of vapor phase)
- H [J/mol] Enthalpy
- S [J/mol/K] Entropy

Corrected means that the Peneloux correction for the density was taken into account.

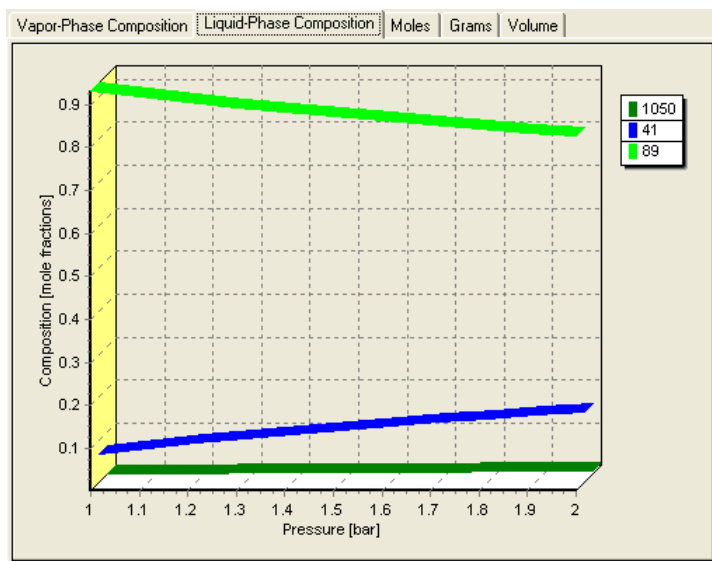
The plot output contains five different charts:



Vapor phase composition.

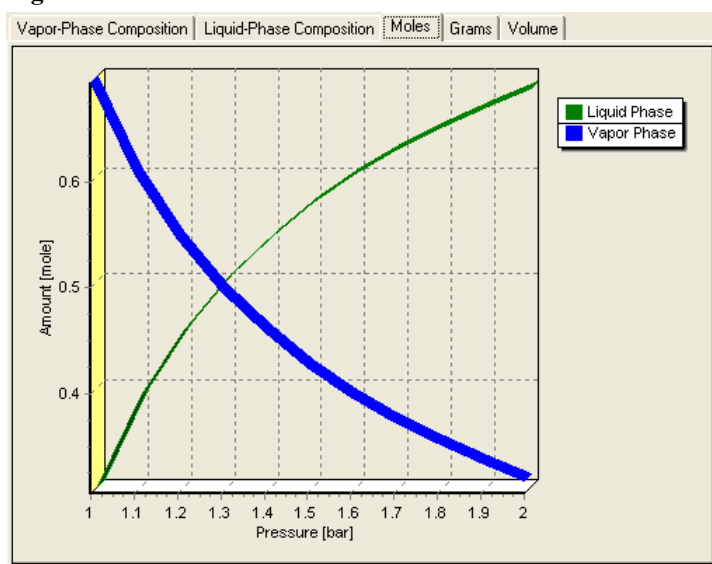
It can be seen that with increasing pressures the fraction of higher boiling compounds gets lower and the fraction of low-boiling substances gets higher (1050 is Carbon Dioxide, 41 Butane, 89 Hexane).

Figure 10



Liquid phase composition

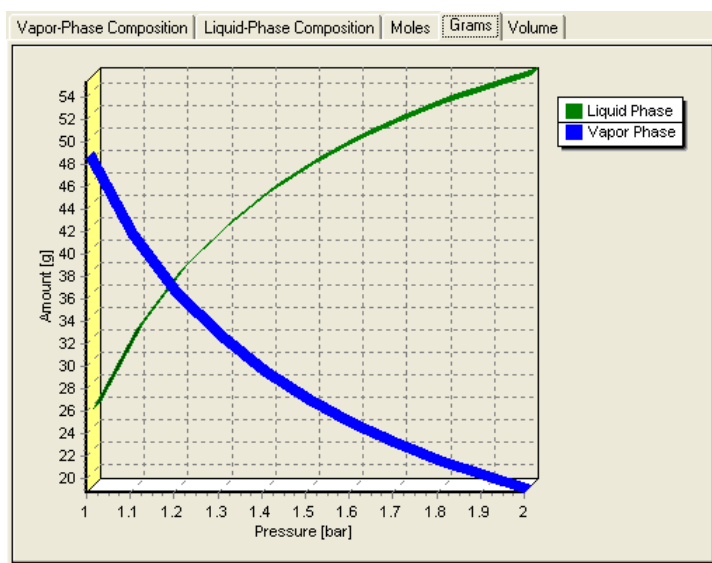
Figure 11



Amount of substance in vapor and liquid phase.

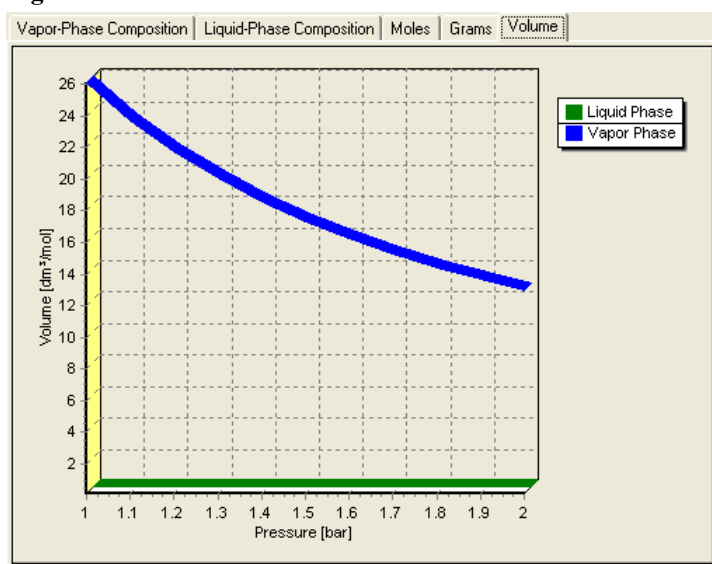
It can be seen that at almost three bar the system has almost no vapor phase anymore.

Figure 12



Mass of substance in liquid and vapor phase. Of course this looks very similar to the moles plot.

Figure 13



Volume of liquid and vapor phase.

The volume vapor phase is of course very impressed by an increased pressure.

Figure 14

Three-Phase Flash Result

The three-phase flash calculation creates only a text output. The output starts with the same lines as the two-phase flash.

The data points for the single temperatures or pressure are a little bit different. The ****Result**** line lists the found phases. In most case this will be liquid and vapor, sometimes only vapor, sometimes only liquid and sometimes liquid-liquid.

```

TEMPERATURE:    323.15 K
PRESSURE       :    1.0000 BAR

**RESULT** 2 PHASES: (LIQUID) (VAPOR)

  I      Z (I)      NZ (I)      X1 (I)      N1 (I)      X2 (I)      N2 (I)
  ---      ---      ---      ---      ---      ---      ---
  1  0.13000  0.13000  0.00159  0.00048  0.18629  0.12952
  2  0.23000  0.23000  0.06697  0.02041  0.30146  0.20959
  3  0.64000  0.64000  0.93144  0.28387  0.51225  0.35613
  ---      ---      ---      ---      ---      ---      ---
N [MOL]          1.00000          0.30476          0.69524
M [G]            74.24           25.67           48.57
V [L/MOL]        0.1327          26.1052
Z=Pv/RT          0.0049          0.9716

TOTAL GIBBS ENERGY OF 2 PHASE MIXTURE:  -2601.4      J

  I  X1*PHI1  X2*PHI2  X2/X1
  ---  ---  ---  ---
  1  .188    .188    118.
  2  .295    .295    4.50
  3  .490    .490    .550
  ---  ---  ---  ---
    
```

Abbreviations - component specific values

- I Component (numbering order from top to bottom as in the Component List)
- Z (I) Total number of moles of component I (normalized to 1 mole of total feed)
- X (I) Mole fraction of component I in the liquid phase
- NZ Number of moles in the liquid and vapor phase phase
- N1 Number of moles in the liquid phase
- N2 Number of moles in the vapor phase phase

Abbreviation – system specific values

- N [MOL] Moles (complete, in liquid phase, in vapor phase)
- M [G] Masses (complete, in liquid phase, in vapor phase)
- V [L/MOL] Volume (of liquid phase, of vapor phase)
- Z=Pv/RT Compressibility (of liquid phase, of vapor phase)

Hint

The three-phase calculation is very slow and unstable and not very reliable.

VLE Calculation

Generate EOS Mixing Rule Parameters allows to predict a VLE data set with EOS mixing rule parameters. These mixing rule parameters might be obtained from new fits or from the library or from a project file.

The grid containing the mixing rule parameters is exactly the same as used in the flash calculation (see section "Flash Calculation" on page 12).

Temperature: 323.15 K Constant Temperature

Pressure: 1 bar Constant Pressure

Start Composition Component 1: 0. Stepwidth x and y: 0.01

Fixed Compositions Component 2 etc.: x2: 0.3

Comp't 1	Comp't 2	a12	b12	c12	a21	b21	c21
41	89	-0.002460237637	-1.727047497e-05	0	-0.002460237637	-1.727047497e-05	0
1050	41	0.1331463213	2.575343567e-05	0	0.1331463213	2.575343567e-05	0
1050	89	0.03713456066	0	0	0.03713456066	0	0

Buttons: Calculate VLE, Show D2P Plot, Show MixPlot, Show MixView

VLE calculations are either isothermal or isobaric and are carried out in one mole percent steps over the entire range from zero mole percent to one hundred mole percent.

For ternary or higher systems it is necessary to specify constant compositions for component 2 and higher.

The calculation yields a table and a plot output:

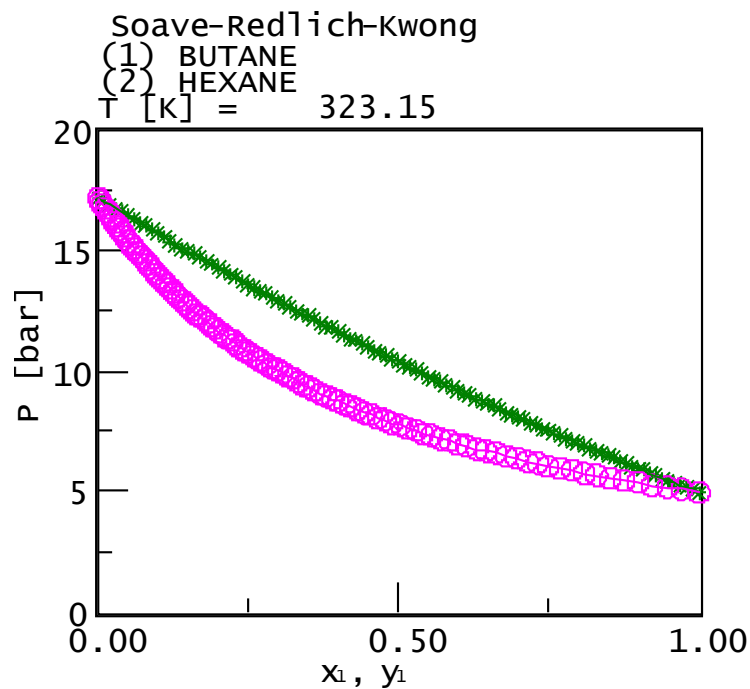
....some lines before this

NO	IT	T [K]	P [BAR]	VV [L/MOL]	VL	X1	X2	Y1	Y2
1	2	323.15	17.194	1.1517	0.1103	0.00000	0.00000	0.00000	0.00000
2	2	323.15	17.039	1.1658	0.1102	0.01000	0.00000	0.00400	0.00000
3	1	323.15	16.886	1.1800	0.1102	0.02000	0.00000	0.00803	0.00000
4	1	323.15	16.733	1.1944	0.1101	0.03000	0.00000	0.01210	0.00000

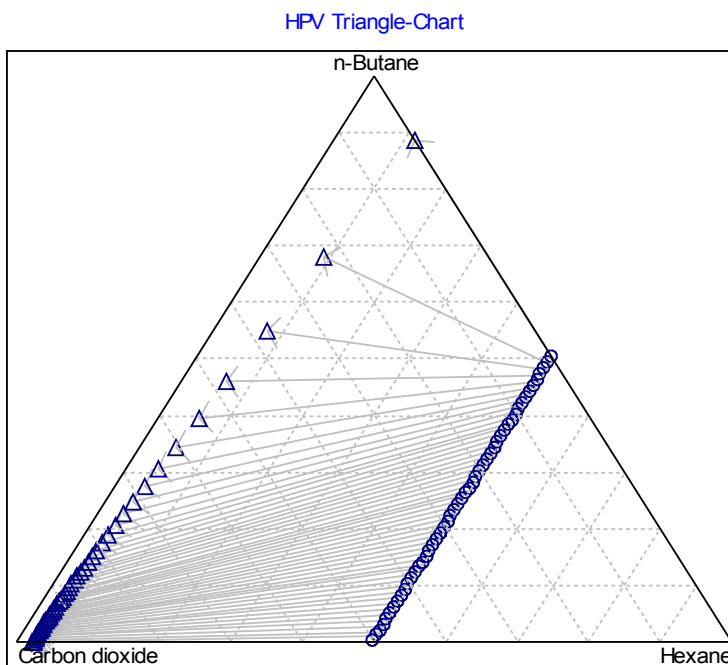
.... some lines after this

The lines in the table contain the temperature, the pressure, the volume of the vapor and the liquid phase and the composition of the liquid and the vapor phase.

The plot (D2P format) shows this table graphically.



A typical result for a ternary system in the MixPlot format shows the constant composition line (Hexane in this example).



Appendix

The EPF File Format

The EPF file is a binary FORTRAN file. It contains number for the equation of state, the mixing rule (for EOS-a and EOS-b), and the g^E model.

After these numbers it contains a list of components with the necessary properties:

1. Critical temperature and pressure
2. Acentric factor
3. Mathias-Copeman parameters (c1, c2, c3)
4. UNIQUAC r and q values
5. A volume at 25 ° for the g^E model Wilson

Behind these pure component properties the list of mixing rule parameters follow.

The first record is ten bytes long and contains five short two-byte numbers:

- Number of components
- Equation of state
 1. Soave-Redlich-Kwong
 2. Peng-Robinson
 3. Redlich-Kwong (currently not supported)
- Mixing rule for EOS a parameter
 1. Quadratic
 2. g^E (Vidal)
 3. g^E (Tochigi)
 4. g^E (Michelsen)
 5. g^E (PSRK)
- Mixing rule for EOS b parameter (always 1)
- g^E model
 1. NRTL
 2. Wilson
 3. UNIQUAC
 4. NRTL (Vidal)
 5. UNIQUAC (res.)

The next record contain the pure component properties. Every record contains one short two-byte number, nine double eight-byte floating point numbers, and 48 characters:

- Component DDB number (short)

- T_c
- P_c
- Acentric factor
- Mathias-Copeman constant
- UNIQUAC r and q values
- A volume for Wilson
- The component name

The next record contain the binary mixing rule parameters. The records contain three double eight-byte floating point numbers and fifty characters:

- G_{ij} and G_{ji}
- α (for NRTL)
- A source definition (up to fifty letters).

Equation of State Formulas

- Soave-Redlich-Kwong

$$P = \frac{RT}{v - b} - \frac{a(T)}{v(v + b)}$$

$$\text{with } \alpha(T) = \left[1 + (0.48 + 1.574\omega - 0.176\omega^2)\right]^2$$

- Peng-Robinson

$$P = \frac{RT}{v - b} - \frac{a(T)}{v(v + b) + b(v - b)}$$

$$\text{with } a(T) = a\alpha(T) \text{ and } \alpha(T) = \left[1 + (0.37464 + 1.54226\omega - 0.26992\omega^2)(1 - T_r^{0.5})\right]^2$$

ω is the acentric factor.

Mixing Rule Equations

- Quadratic Mixing Rule

$$a = \sum_i \sum_j y_i y_j a_{ij}$$

$$\text{with } a_{ij} = (a_{ii} a_{jj})^{0.5} (1 - k_{ij})$$

- g^E Mixing Rule Vidal (Huron/Vidal)

$$\frac{\bar{a}^{-E}}{RT} = A_0 \left[\frac{a}{bRT} - \sum x_i \frac{a_{ii}}{b_i RT} \right] = \frac{g_\infty^E}{RT}$$

$$A_0 = \ln \frac{v}{v + b} = \ln \frac{1}{2}$$

- g^E Mixing Rule Michelsen

$$\frac{\bar{a}^{-E}}{RT} = A_0 \left[\frac{a}{bRT} - \sum x_i \frac{a_{ii}}{b_i RT} \right] = \frac{g_0^E}{RT} + \sum x_i \ln \frac{b}{b_i} \quad A_0 = \ln \frac{v}{v+b} = \ln \frac{1.235}{2.235} = -0.593$$

- g^E Mixing Rule PSRK

$$\frac{\bar{a}^{-E}}{RT} = A_0 \left[\frac{a}{bRT} - \sum x_i \frac{a_{ii}}{b_i RT} \right] = \frac{g_0^E}{RT} + \sum x_i \ln \frac{b}{b_i} \quad A_0 = \ln \frac{v}{v+b} = \ln \frac{1.1}{2.1} = -0.64663$$

- Panagiotopoulos/Reid

$$a = \sum_i \sum_j y_i y_j a_{ij}$$

$$\text{with } a_{ij} = \sqrt{a_i a_j} \left[1 - k_{ij} + (k_{ij} - k_{ji}) x_i \right]$$